

Acs Physical Chemistry Study Guide

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$\ln K = \frac{\Delta_r S^\circ}{R} - \frac{\Delta_r H^\circ}{RT}$ The equilibrium for the dissociation of one mole of iodine vapour is 3.94×10^{-4} at 900K and the standard enthalpy is 154 kJmol⁻¹. Calculate the equilibrium constant at 1200K assuming the standard enthalpy remains constant. $\ln(K_2/K_1) = -\frac{\Delta_r H^\circ}{R} \left(\frac{1}{T_2} - \frac{1}{T_1}\right)$, ANS = 6.7×10^{-2} at 1200K.

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Physical chemistry is the study of how matter behaves on a molecular and atomic level and how chemical reactions occur. Based on their analyses, physical chemists may develop new theories, such as how complex structures are formed. Physical chemists often work closely with materials scientists to research and develop potential uses for new materials.

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